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| **STEP 1** | Find the Fluorine on the Periodic Table |
| **STEP 2** | Use the Atomic Number to figure out how many protons and electrons are in F (Hint: use PEN)  ANSWER = \_\_\_\_ |
| **STEP 3** | Use the Atomic Mass to figure out how many neutrons are in F (Hint: Neutrons =Atomic Number - Atomic Mass) ANSWER = \_\_\_ -\_\_\_    =\_\_\_\_\_ |
| **STEP 4** | Use the Octet Rule to calculate how many shells (orbits) F has (Hint: you can have 2 e- in shell 1, 8 e- in shell 2) ANSWER = \_\_\_\_ |
| **STEP 5** | Draw a circle for the nucleus and write how many p, n are in it |
| **STEP 6** | Draw shells and add dots for each electron |

STEPS FOR DRAWING A BOHR-RUTHERFORD DIAGRAM

FLUORINE BR EXAMPLE: Boron-11

